**(BT)**

# BIO TECHNOLOGY
## *INSTRUCTIONS TO CANDIDATES*

1. Candidates should write their Hall Ticket Number only in the space provided at the top left hand corner of this page, on the leaflet attached to this booklet and also in the space provided on the OMR Response Sheet. **BESIDES WRITING, THE CANDIDATE SHOULD ENSURE THAT THE APPROPRIATE CIRCLES PROVIDED FOR THE HALL TICKET NUMBERS ARE SHADED USING H.B. PENCIL ONLY ON THE OMR RESPONSE SHEET. DO NOT WRITE HALL TICKET NUMBER ANY WHERE ELSE.**

2. Immediately on opening this Question Paper Booklet, check:
   (a) Whether **200** multiple choice questions are printed (**50** questions in Mathematics, **25** questions in Physics, **25** questions in Chemistry and **100** questions in Engineering)
   (b) In case of any discrepancy immediately exchange the Question paper Booklet of same code by bringing the error to the notice of invigilator.

3. Use of Calculators, Mathematical Tables and Log books is not permitted.

4. **Candidate must ensure that he/she has received the Correct Question Booklet, corresponding to his/her branch of Engineering.**

5. **Candidate should ensure that the booklet Code and the Booklet Serial Number, as it appears on this page is entered at the appropriate place on the OMR Response Sheet by shading the appropriate circles provided therein using H.B. pencil only. Candidate should note that if they fail to enter the Booklet Serial Number and the Booklet Code on the OMR Response Sheet, their Answer Sheet will not be valued.**

6. **Candidate shall shade one of the circles 1, 2, 3 or 4 corresponding question on the OMR Response Sheet using H.B. Pencil only. Candidate should note that their OMR Response Sheet will be invalidated if the circles against the question are shaded using Black / Blue ink pen / Ball pen / any other pencil other than H.B. Pencil or if more than one circle is shaded against any question.**

7. One mark will be awarded for every correct answer. **There are no negative marks.**

8. The OMR Response Sheet will not be valued if the candidate :
   (a) Writes the Hall Ticket Number in any part of the OMR Response Sheet except in the space provided for the purpose.
   (b) Writes any irrelevant matter including religious symbols, words, prayers or any communication whatsoever in any part of the OMR Response Sheet.
   (c) Adopts any other malpractice.

9. Rough work should be done only in the space provided in the Question Paper Booklet.

10. No loose sheets or papers will be allowed in the examination hall.

11. Timings of Test: 10.00 A.M. to 1.00 P.M.

12. Candidate should ensure that he / she enters his / her name and appends signature on the Question paper booklet, leaflet attached to this question paper booklet and also on the OMR Response Sheet in the space provided. Candidate should ensure that the invigilator puts his signature on this question paper booklet, leaflet attached to the question paper booklet and also on the OMR Response Sheet.

13. Before leaving the examination hall candidate should **return both the OMR Response Sheet and the leaflet attached to this question paper booklet** to the invigilator. Failure to return any of the above shall be construed as malpractice in the examination. **Question paper booklet may be retained by the candidate.**

14. This booklet contains a total of **32** pages including Cover page and the pages for Rough Work.

*Note:* (1) Answer **all** questions.

(2) *Each question carries **1** mark. There are no negative marks.*

(3) *Answer to the questions must be entered only on OMR Response Sheet provided separately by completely shading with H.B. Pencil, only one of the circles 1, 2, 3 or 4 provided against each question, and which is most appropriate to the question.*

## MATHEMATICS

1. If $A = \begin{bmatrix} 3 & 0 & 0 \\ 0 & 3 & 0 \\ 0 & 0 & 3 \end{bmatrix}$, then $A^4 =$

(1) $3I$      (2) $9I$      (3) $27I$      (4) $81I$

2. If $A = \begin{bmatrix} 0 & 2 & 1 \\ -2 & 0 & -2 \\ -1 & x & 0 \end{bmatrix}$ is a skew symmetric matrix, then the value of $x$ is

(1) 1      (2) 2      (3) 3      (4) 4

3. What is the number of all possible matrices with each entry as 0 or 1 if the order of matrices is $3 \times 3$

(1) 64      (2) 268      (3) 512      (4) 256

4. If $A = \begin{bmatrix} 1 & i & -i \\ i & -i & 1 \\ -i & 1 & i \end{bmatrix}$, then $|A| =$

(1) 1      (2) 2      (3) 3      (4) 4

5. The solution of a system of linear equations $2x - y + 3z = 9, x + y + z = 6, x - y + z = 2$ is

(1) $x = -1, y = -2, z = -3$
(2) $x = 3, y = 2, z = 1$
(3) $x = 2, y = 1, z = 3$
(4) $x = 1, y = 2, z = 3$

6. If $\dfrac{1}{x^2 + a^2} = \dfrac{A}{x + ai} + \dfrac{B}{x - ai}$ then A = _____, B = _____.

(1) $\dfrac{1}{2ai}, -\dfrac{1}{2ai}$
(2) $-\dfrac{1}{2ai}, \dfrac{1}{2ai}$
(3) $\dfrac{1}{ai}, -\dfrac{1}{ai}$
(4) $-\dfrac{1}{ai}, \dfrac{1}{ai}$

7. If $\dfrac{2x + 4}{(x-1)^3} = \dfrac{A_1}{(x-1)} + \dfrac{A_2}{(x-1)^2} + \dfrac{A_3}{(x-1)^3}$ then $\displaystyle\sum_{i=1}^{3} A_i$ is equal to

(1) $A_2$
(2) $2A_2$
(3) $4A_2$
(4) $4A_1$

8. The period of the function $f(x) = |\sin x|$ is

(1) $\pi$
(2) $2\pi$
(3) $3\pi$
(4) $4\pi$

9. If A+B=45°, then $(1 - \cot A) \cdot (1 - \cot B)$ is

(1) 1
(2) 0
(3) 2
(4) −1

10. The value of sin 78° + cos 132° is

(1) $\dfrac{\sqrt{5} + 1}{4}$
(2) $\dfrac{\sqrt{5} + 1}{2}$
(3) $\dfrac{\sqrt{5} - 1}{2}$
(4) $\dfrac{\sqrt{5} - 1}{4}$

11. If A+B+C = $\pi$, then sin2A + sin2B + sin2C =

(1) 4 cosA sinB cosC
(2) 4 sinA cosB sinC
(3) 4 cosA cosB cosC
(4) 4 sinA sinB sinC

12. The principal solution of Tan$x$ = 0 is

(1) $x = n\pi, n \in Z$
(2) $x = 0$
(3) $x = (2n+1) \pi/2, n \in Z$
(4) $x = n\pi + \alpha, n \in Z$

13. The value of $\text{Tan}^{-1}(2) + \text{Tan}^{-1}(3)$ is

(1) $\dfrac{\pi}{4}$      (2) $\dfrac{\pi}{2}$      (3) $\dfrac{\pi}{3}$      (4) $\dfrac{3\pi}{4}$

14. If the sides of a right angle triangle are in A.P., then the ratio of its sides is

(1) 1:2:3      (2) 2:3:4      (3) 3:4:5      (4) 4:5:6

15. The value of $r.r_1.r_2.r_3$ is

(1) $\Delta^2$      (2) $\Delta^{-2}$      (3) $\Delta^{-3}$      (4) $\Delta^4$

16. $\dfrac{1}{r1}+\dfrac{1}{r2}+\dfrac{1}{r3}=$

(1) $\dfrac{1}{r}$      (2) $\dfrac{1}{2r}$      (3) $\dfrac{1}{R}$      (4) $\dfrac{1}{\Delta}$

17. If $a=6, b=5, c=9$, then the value of angle A is

(1) $\cos^{-1}(2/9)$      (2) $\cos^{-1}(2/5)$      (3) $\cos^{-1}(7/9)$      (4) $\cos^{-1}(1/3)$

18. The polar form of complex number $1-i$ is

(1) $\sqrt{2}\,e^{-i\pi/4}$      (2) $\sqrt{2}\,e^{i\pi/4}$      (3) $\sqrt{2}\,e^{i\pi/2}$      (4) $\sqrt{2}\,e^{-i\pi/2}$

19. If $1, \omega, \omega^2$ be the cube roots of unity, then the value of $2^{\omega^3}.2^{\omega^5}.2^{\omega}$ is

(1) $\omega$      (2) $\omega^2$      (3) 1      (4) 0

20. The intercept made on X-axis by the circle $x^2+y^2+2gx+2fy+c = 0$ is

(1) $\sqrt{g^2-c}$      (2) $\sqrt{f^2-c}$      (3) $2.\sqrt{g^2-c}$      (4) $2.\sqrt{f^2-c}$

21. If one end of the diameter of the circle $x^2+y^2-5x-8y+13 = 0$ is (2, 7), then the other end of the diameter is

(1) (3, 1)      (2) (1, 3)      (3) (−3, −1)      (4) (−1, −3)

22. The radius of the circle $\sqrt{1+m^2}(x^2+y^2)-2cx-2mcy=0$ is
   (1) $2c$　　　　(2) $4c$　　　　(3) $c/2$　　　　(4) $c$

23. The parametric equations of the ellipse $\dfrac{x^2}{a^2}+\dfrac{y^2}{b^2}=1$ are

   (1) $x = a\sec\theta,\ y = b\tan\theta$　　　　(2) $x = b\sin\theta,\ y = a\cos\theta$
   (3) $x = a\cos\theta,\ y = b\sin\theta$　　　　(4) $x = a\csc\theta,\ y = b\cot\theta$

24. The equation of the directrix of the parabola $2x^2=-7y$ is
   (1) $8y+7=0$　　(2) $8y-7=0$　　(3) $7y+8=0$　　(4) $8x-7=0$

25. The condition for a straight line $y = mx+c$ to be a tangent to the hyperbola $\dfrac{x^2}{a^2}-\dfrac{y^2}{b^2}=1$ is

   (1) $c = a/m$　　(2) $c^2 = a^2m^2 - b^2$　　(3) $c^2 = a^2m^2 + b^2$　　(4) $c^2 = a/m$

26. $\underset{x\to 1}{Lt}\ \dfrac{\sqrt{5x-4}-\sqrt{x}}{x-1}$ is

   (1) 3　　　　(2) 2　　　　(3) 4　　　　(4) 1

27. $\log i =$
   (1) $\pi/2$　　　　(2) $\pi/4$　　　　(3) $i\pi/2$　　　　(4) $i\pi/4$

28. $\dfrac{d}{dx}\left[\log_7 X\right]=$

   (1) $\dfrac{1}{x}$　　(2) $X\log_7^e$　　(3) $\dfrac{1}{x}\log^7_e$　　(4) $\dfrac{1}{x}\log_7^e$

29. $\dfrac{d}{dx}\left[2\cosh x\right]=$

   (1) $\dfrac{e^x+e^{-x}}{2}$　　(2) $\dfrac{e^x-e^{-x}}{2}$　　(3) $e^x+e^{-x}$　　(4) $e^x-e^{-x}$

30. $\dfrac{d}{dx}\left[\cos^{-1}\left(\dfrac{1-x^2}{1+x^2}\right)\right]=$

(1) $\dfrac{1}{1+x^2}$    (2) $\dfrac{-1}{1+x^2}$    (3) $\dfrac{2}{1+x^2}$    (4) $\dfrac{-2}{1+x^2}$

31. If $x=at^2, y=2at$, then $\dfrac{dy}{dx}=$

(1) $\sqrt{\dfrac{y}{x}}$    (2) $\sqrt{\dfrac{x}{a}}$    (3) $\sqrt{\dfrac{a}{x}}$    (4) $\sqrt{\dfrac{x}{y}}$

32. The derivative of $e^x$ with respect to $\sqrt{x}$ is

(1) $\dfrac{2\sqrt{x}}{e^x}$    (2) $2\sqrt{x}\,e^x$    (3) $\dfrac{e^x}{2\sqrt{x}}$    (4) $\sqrt{x}.e^x$

33. The equation of the normal to the curve $y=5x^4$ at the point $(1, 5)$ is

(1) $x+20y=99$    (2) $x+20y=101$    (3) $x-20y=99$    (4) $x-20y=101$

34. The angle between the curves $y^2=4x$ and $x^2+y^2=5$ is

(1) $\dfrac{\pi}{4}$    (2) $\tan^{-1}(2)$    (3) $\tan^{-1}(3)$    (4) $\tan^{-1}(4)$

35. If $u=x^3y^3$ then $\dfrac{\partial^3 u}{\partial x^3}+\dfrac{\partial^3 u}{\partial y^3}=$

(1) $6(x^3+y^3)$    (2) $6x^3y^3$    (3) $6x^3$    (4) $6y^3$

36. $\int\operatorname{cosec} x\,dx=$

(1) $\log(\operatorname{cosec} x + \cot x)+C$    (2) $\log(\cot x/2)+C$
(3) $\log(\tan x/2)+C$    (4) $-\operatorname{cosec} x.\cot x+C$

# CHEMISTRY

76. The valency electronic configuration of Phosphorous atom (At.No. 15) is
    (1) $3s^2 3p^3$      (2) $3s^1 3p^3 3d^1$      (3) $3s^2 3p^2 3d^1$      (4) $3s^1 3p^2 3d^2$

77. An element 'A' of At.No.12 combines with an element 'B' of At.No.17. The compound formed is
    (1) covalent AB      (2) ionic $AB_2$      (3) covalent $AB_2$      (4) ionic AB

78. The number of neutrons present in the atom of $_{56}Ba^{137}$ is
    (1) 56      (2) 137      (3) 193      (4) 81

79. Hydrogen bonding in water molecule is responsible for
    (1) decrease in its freezing point      (2) increase in its degree of ionization
    (3) increase in its boiling point      (4) decrease in its boiling point

80. In the HCl molecule, the bonding between hydrogen and chlorine is
    (1) purely covalent (2) purely ionic      (3) polar covalent (4) complex coordinate

81. Potassium metal and potassium ions
    (1) both react with water      (2) have the same number of protons
    (3) both react with chlorine gas      (4) have the same electronic configuration

82. 5.85 gms of sodium chloride were dissolved in water and the solution made upto 100 ml in a standard flask. 10 ml of this solution were pipetted out into another flask and made up with distilled water into 100 ml of solution. The concentration of the sodium chloride solution now is
    (1) 0.1 M      (2) 1.0 M      (3) 0.5 M      (4) 0.25 M

83. Concentration of a 1.0 M solution of phosphoric acid in water is
    (1) 0.33 N      (2) 1.0 N      (3) 2.0 N      (4) 3.0 N

84. Which of the following is a Lewis acid?
    (1) Ammonia      (2) Berylium chloride
    (3) Boron trifluoride      (4) Magnesium oxide

85. Which of the following constitutes the components of a buffer solution?
    (1) Potassium chloride and potassium hydroxide
    (2) Sodium acetate and acetic acid
    (3) Magnesium sulphate and sulphuric acid
    (4) Calcium chloride and calcium acetate

86. Which of the following is an electrolyte?
    (1) Acetic acid    (2) Glucose    (3) Urea    (4) Pyridine

87. Calculate the Standard emf of the cell, $Cd/Cd^{+2}//Cu^{+2}/Cu$ given that $E^0\ Cd/Cd^{+2} = 0.44V$ and $E^0\ Cu/Cu^{+2} = (-)\ 0.34\ V$.
    (1) (–) 1.0 V    (2) 1.0 V    (3) (–) 0.78 V    (4) 0.78 V

88. A solution of nickel chloride was electrolysed using Platinum electrodes. After electrolysis,
    (1) nickel will be deposited on the anode    (2) $Cl_2$ gas will be liberated at the cathode
    (3) $H_2$ gas will be liberated at the anode    (4) nickel will be deposited on the cathode

89. Which of the following metals will undergo oxidation fastest?
    (1) Cu    (2) Li    (3) Zinc    (4) Iron

90. Which of the following cannot be used for the sterilization of drinking water?
    (1) Ozone    (2) Calcium Oxychloride
    (3) Potassium Chloride    (4) Chlorine water

91. A water sample showed it to contain 1.20 mg/litre of magnesium sulphate. Then, its hardness in terms of calcium carbonate equivalent is
    (1) 1.0 ppm    (2) 1.20 ppm    (3) 0.60 ppm    (4) 2.40 ppm

92. Soda used in the L-S process for softening of water is, Chemically.
    (1) sodium bicarbonate    (2) sodium carbonate decahydrate
    (3) sodium carbonate    (4) sodium hydroxide (40%)

93. The process of cementation with zinc powder is known as
    (1) sherardizing    (2) zincing    (3) metal cladding    (4) electroplating